

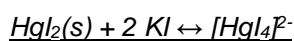
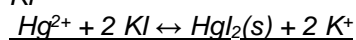
Pregled reakcija kationa 2. skupine

Reagens	Pb^{2+}	Hg^{2+}	Cu^{2+}	Bi^{3+}	Cd^{2+}	Sb^{3+}	Sn^{2+}
H_2S							
NaOH (ekv.)		$Hg^{2+} + 2 NaOH \leftrightarrow HgO(s) + H_2O + 2 Na^+$					
NaOH (suv.)	$Pb(OH)_2(s) + 2NaOH \leftrightarrow PbO_2^{2-} + 2 H_2O + 2 Na^+$	$Hg^{2+} + 2 NaOH \leftrightarrow HgO(s) + H_2O + 2 Na^+$				$Sb(OH)_3(s) + 2 NaOH \leftrightarrow SbO_3^{3-} + 2 H_2O + 2 Na^+$	$Sn(OH)_2(s) + 2 NaOH \leftrightarrow SnO_2^{2-} + 2 H_2O + 2 Na^+$
NH_4OH (ekv.)		$Hg^{2+} + 2 Cl^- + NH_3 \leftrightarrow HgNH_2Cl(s) + HCl$	$Cu(OH)_2(s) + 4 NH_3 \leftrightarrow [Cu(NH_3)_4]^{2+} + 2 OH^-$				
NH_4OH (suv.)		$Hg^{2+} + 2 Cl^- + NH_3 \leftrightarrow HgNH_2Cl(s) + HCl$	$Cu(OH)_2(s) + 4 NH_3 \leftrightarrow [Cu(NH_3)_4]^{2+} + 2 OH^-$		$Cd(OH)_2(s) + 4 NH_3 \leftrightarrow [Cd(NH_3)_4]^{2+} + 2 OH^-$		

Karakteristične reakcije kationa 2. skupine

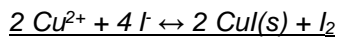
1. Živa, Hg^{2+}

a. KI

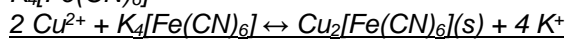


2. Bakar, Cu^{2+}

a. KI



b. $K_4[Fe(CN)_6]$



3. Olovo, Pb^{2+}

a. KI

b. K_2CrO_4
